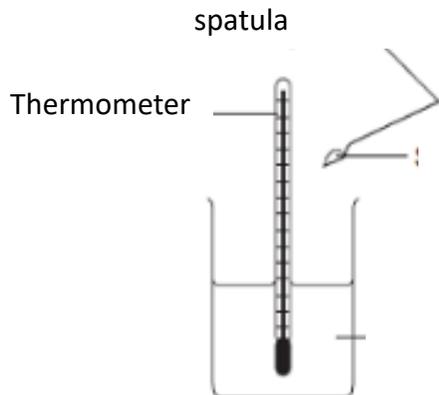


Key Word	Definition
Catalyst	Substances that speed up chemical reactions but are unchanged at the end.
Exothermic reaction	One in which energy is given out, usually as heat or light.
Endothermic reaction	One in which energy is taken in, usually as heat.
Chemical bond	Force that holds atoms together in molecules.
Activation energy	The minimum(lowest) amount of energy needed for a reaction to take place



Investigating energy changes. A raise in temperature will mean the reaction is exothermic and a decrease in temperature will mean the reaction is endothermic.

Uses of catalysts

- Metals such as platinum and palladium are used in car exhausts to make gases less harmful
- Nickel is used to make margarine
- Enzymes are biological catalysts

Examples of exothermic reactions -Combustion (burning), neutralisation, respiration

Examples of endothermic reactions -Photosynthesis, Ice packs

